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IB Chemistry Topic 2 Atomic structure

2.1 The nuclear atom ~~IB Chemistry~~

~~Topic 2 Atomic structure 12.1~~

~~Electrons in atoms HL~~ Topic 12 - HL

Atomic Structure NS-IB Unit 04 Atomic

Structure Slides 47-51 IB Chemistry

Topic 2 Atomic structure 2.2 Electron

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configuration IB Chemistry: Atomic
Structure Overview 2.1/2.2 Atomic
structure (SL) Atomic Structure - IB
Chemistry SL

Atomic Structure \u0026amp; Electrons -
Inside Atoms - Neutrons Protons
Electrons #VCE#IB#AP#A level
Chem1.3IB Chemistry HL Topic 12:

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Atomic theory (part 1) Atomic structure
and Mass Spectroscopy (IB Chemistry
2.1) IB Chemistry SL/HL Topic 2:
Pearson (2014) Textbook Practice
Questions ~~IB EXAM RESULTS~~
~~REACTION!! [May 2018 Session]~~
~~Katie Tracy~~ HOW TO STUDY FOR
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Chemistry: ~~Order of Filling Orbitals~~
(~~Atomic Structure~~) 2.2 Hydrogen
emission spectrum (SL)~~

Energy levels, sublevels, \u0026amp; orbitals

12.1 Calculating ionisation energy

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(new) (HL) How Small Is An Atom?

Spoiler: Very Small. HL IB Chemistry

31 Mistakes 12.1 Limit of

Convergence and Calculations [HL IB

Chemistry] 5 Minute Lesson:

CHEMISTRY #1: Atomic Structure IB

Chemistry □ Topic 2 Atomic structure □

Find the number of p, n, e, relative

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~~Atomic mass Chemistry — Atomic
Structure — EXPLAINED!~~ The whole of
ATOMIC STRUCTURE in 20 minutes!
AQA C1 GCSE 9-1 Combined Science
or Chemistry Revision Atomic
Structure | A-level Chemistry | OCR,
AQA, Edexcel Science Class 7 Unit 6
\"Structure of Atom\" Page 57 Topic

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Structure of Atom IB Chemistry Topic
1 Stoichiometric relationships Topic
1.1 Introduction to Chemistry SL
~~Atomic Number, Atomic Mass, and the
Atomic Structure | How to Pass
Chemistry Science | Prep.1 | Atomic
structure of matter | Part (2/4) | Unit
One - Lesson Three Unit Atomic~~

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Atomic Mass Units (AMU): 1/12th of the mass of a carbon ^{12}C atom in its ground state. This is used to express masses of atomic particles. 1 AMU =

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1.6605402 x 10⁻²⁷ kg 2.1 The nuclear
atom □ IB Alchemy Completion of the
Group 4

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The atomic mass scale is based on

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the mass of one atom of the carbon 12 isotope (6 protons and 6 neutrons) being equal to exactly 12 units - all other masses are compared to this. For example the relative atomic mass of helium is 4 which has been calculated from the fact that it is one third as heavy as carbon 12.

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the mass of a carbon ^{12}C atom in its
ground state. This is used to express
masses of atomic particles. 1 AMU =

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1.6605402 x 10⁻²⁷ kg 2.1 The nuclear
atom □ IB Alchemy Created Date:
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Atomic Structure and Nuclear Page
6/23

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Unit Atomic Structure Ib Expectations
Assessment Criteria Atomic Mass

Units (AMU): 1/12th of the mass of a
carbon ^{12}C atom in its ground state.

This is used to express masses of
atomic particles. $1 \text{ AMU} = 1.6605402 \times 10^{-27} \text{ kg}$

2.1 The nuclear atom IB
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Expectations Atomic Mass Units
(AMU): $\frac{1}{12}$ th of the mass of a carbon

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12 atom in its ground state. This is used to express masses of atomic particles. $1 \text{ AMU} = 1.6605402 \times 10^{-27} \text{ kg}$

2.1 The nuclear atom

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state. This is used to express masses
of atomic particles. $1 \text{ AMU} =$
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Criteria Topic 2: Atomic structure.
Notes for the Core IB Chemistry
module: Topic 2: Atomic structure.

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These have been made according to
the Page 5/28

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Structure Ib Expectations Assessment
Criteria Unit Atomic Structure Ib
Expectations Atomic Mass Units

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(AMU): 1/12th of the mass of a carbon
-12 atom in its ground state. This is
used to express masses of atomic
particles. $1 \text{ AMU} = 1.6605402 \times 10^{-27} \text{ kg}$

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Apollo 9 was a March 1969 human
spaceflight, the third in NASA's Apollo
program. Flown in low Earth orbit, it

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was the second crewed Apollo mission that the United States launched via a Saturn V rocket, and was the first flight of the full Apollo

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IB Chemistry HL-II Summer Review

Unit 1 □ Atomic Structure IB 2.1 The
nuclear atom 1. State the number of
protons, neutrons, and electrons in

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each of the following: a. ^{65}Cu b. $^{15}\text{N}_3$ -c. $^{137}\text{Ba}^{2+}$ 2. Determine the relative atomic mass of copper (to 2 decimal places) given the following natural abundances: ^{63}Cu 76.00% and ^{65}Cu 24.00% 3.

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